	(08 Version C
	Name Key
	HONORS CHEMISTRY QUIZ - MEASUREMENTS AND MATH
(lea)	MULTIPLE CHOICE - Write the letter of the best possible answer to the left of the number.
	a(n)
	A. hibiscus B. concave lid C. depression zone D. meniscus
	DA 2. A negative exponent written with a number in scientific notation means
	A. the number is less than one
	B. the number is greater than one C. move the decimal to the right for standard notation
	D. the decimal was moved to the left when it was put into standard notation
	3. When adding and subtracting with significant figures the answer must have
	A. the same number of significant figures as the number with the least significant figures B. the same number of significant figures as the number with the most significant figures
	C. the same number of decimal places as the number with the least decimal places
	D. the same number of decimal places as the number with the most decimal places
	B 4. The standard unit for volume is the
	A. Angstrom B. liter C. kilogram D. meter
	5. When multiplying and dividing with significant figures the answer must have the same significant figures as the number with
	A. the most significant figures C. the least decimal places
	B. the most decimal places D. the least significant figures
(200	SIGNIFICANT FIGURES: How many significant figures are in each of the measurements below. Circle
	the significant figures.
	6. 0.302000 g 6 10. 0.03407 g 4
	7. 30040 g 4 11. (22000 g 2

9.3.0x 10 ⁻⁴ g 2

8. 0.0000**9**g ___l

12. (3.0)g 2



CALCULATIONS: Your answer should contain the correct significant figures and the correct units.

-1 curts

13.
$$24,456g + 134g + 20.1g + 0.55g = 24611g$$
 (nod.p.)

- I mat

14.
$$2.96 L \div (4.0 \times 10^2) s = -0074 L/S$$

-12

15. $(3007 \text{ cm}) (0.004 \text{ cm}) (9.6 \times 10^{1} \text{ cm}) = 1000 \text{ cm}^3 \text{ or } 1 \times 10^3 \text{ cm}^3$

(2,

<u>SCIENTIFIC NOTATION</u>: Convert the scientific notation to decimal form and the decimals to scientific notation. Show correct significant figures.

-lexp

16. 4.032 x 10 ⁵	403200	
17. 50140	5.014 x 104	
18. 0.0034002	3.4002 x 10 -3	
19. 2.36×10^{-3}	.00236	

<u>DIMENSIONAL ANALYSIS:</u> Solve the following problems. Use unit analysis if appropriate. Use correct significant figures and units. <u>SHOW ALL WORK and circle your final answer.</u>

(8)

20. Convert 0.687 g/cm3 to lb/gal.

-1 of	.68791	1(16)	1 cm3	1000mt	14	4 gt	(Ih)
-1 mata -1 conv	1 cms	454.0	Int	14 1	1.06gt	Igal =	5.71 b

-loy 7

21. The diameter of an atom is 1.95 Angstroms. When 5.5×10^9 atoms are laid side by side, what is the length of the row in feet? (An angstrom is a metric unit of length and is abbreviated with an A. $1A = 1 \times 10^{-8}$ cm)

Facts:



3 22. Mrs. Hugg....

1.00 mile. One of Mrs. Fugg....

walk to burn the calories from 3 peanut butter.

Facts: | Sand | 130 cal | 900 m
| 980 cal | Imi | 1 step

2 44

3 sand | 980 cat | Imi | 1 lolkm | 1000 m | 1 step = 4.0 × 10 steps

1 | 1 sand | 130 cat | Imi | 1 km | 900 m | 40,000 steps

2 44

40,000 steps

70

400

300

100 g

500 g

Reading Scales. Record the measurement of each scale. Include units.

1 number 23. Thermometer 37.0°C

man oft

- 24. Balance 480.50 g
- 25. Graduated Cylinder 40.0 mL

200

-	-									
lun.	ann him	hmlan	malan	milan	hilan	milan	milno	milan	hulan	limil
0	1	2	3	4	5	6	7	8	9	10 g



