

81 pass pts.  
73

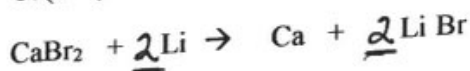
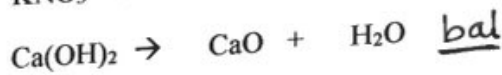
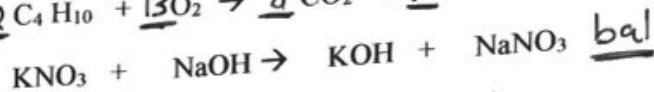
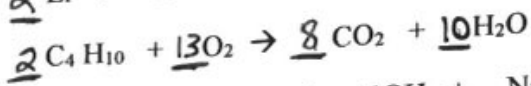
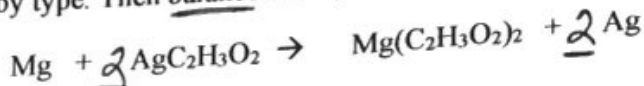
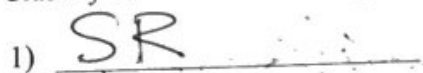
# Honors Chemistry Quiz - Types of Reactions and Predicting Products

Recovery Quiz

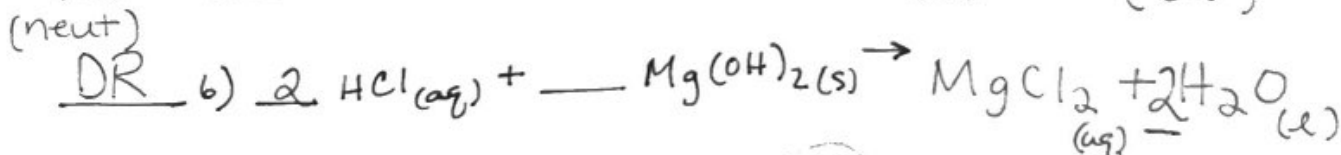
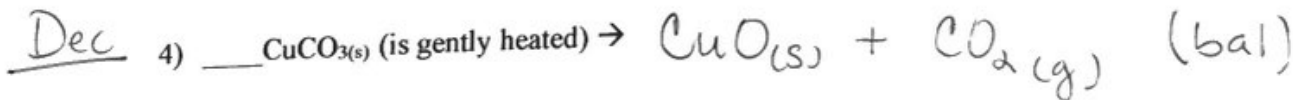
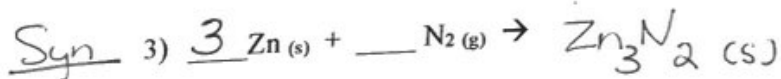
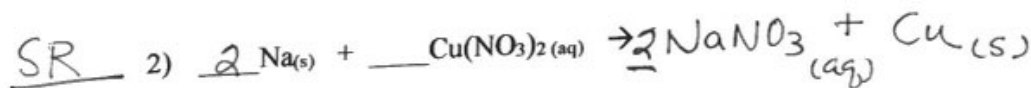
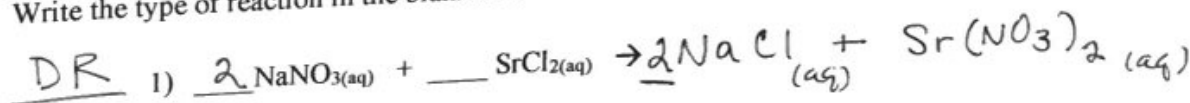
Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

D = decomposition, S = synthesis, SR = single replacement, DR = double replacement, C = combustion

1. Classify each of the following reactions by type. Then balance the equation.



2. Complete the reaction by predicting the products. Include states of matter. Then balance!  
Write the type of reaction in the blank to the left.

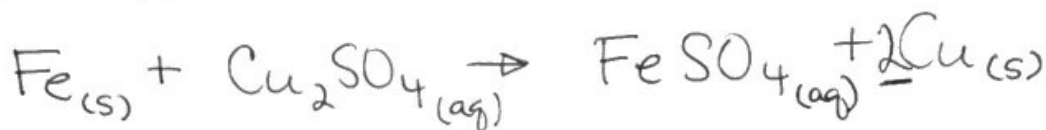


(1) type  
(1) bal

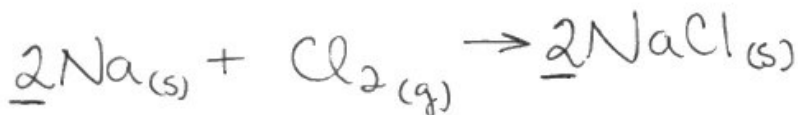
(1) type  
(2) prod  
(1) states  
(1) bal

- 24
3. Write balanced chemical equations for the reactions below. Follow these steps:
- (2) 1. Write good formulas for the reactants.
  - (1) 2. Figure out what type of reaction (write in blank at the left) and then use that information to predict the products.
  - (2) 3. Write good formulas for the products.
  - (1) 4. Balance the equation.
  - (1) 5. Write the phase for each reactant and product.

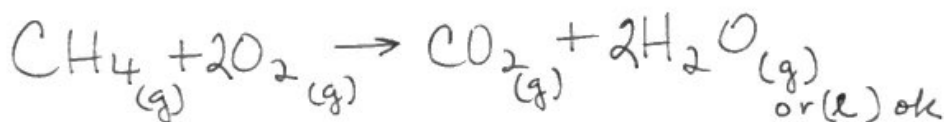
SR 1) solid iron (II) reacts with a solution of copper (I) sulfate



Syn 2) solid sodium reacts with chlorine gas



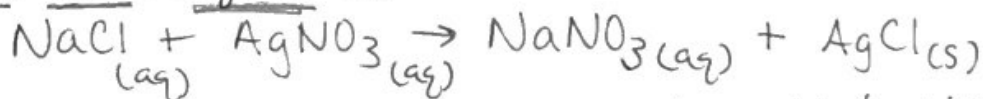
\* C 3) the combustion of methane <sup>gas</sup> (CH<sub>4</sub>)



\* 4) 4. What are the four signs that a chemical reaction has occurred?

1. ppt forms
2. gas forms
3. color Δ
4. heat/light prod (Temp change)

6) 5. When aqueous NaCl reacts with aqueous AgNO<sub>3</sub>, what are the spectator ions? What is the net ionic equation?



spectator ions: Na<sup>+</sup>, NO<sub>3</sub><sup>-</sup>

