**Basic Ideas of Quantum Mechanics (continued)**

1. Electrons and atoms can exist only in certain energy states. When an electron changes its energy state, it must emit or absorb just enough energy to bring it to the new energy state.

2. Electrons or atoms emit energy in the form of light. The frequency of the light emitted is related to the energy change.

3. The allowed energy states of electrons or atoms can be described by sets of numbers called \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

4. Pauli Exclusion Principle

5. Aufbau Principle

6. Hund’s Rule

7. Heisenberg Uncertainty Principle